heptylborane (12.1 mmoles), prepared from 1-heptene, 3.92 g (40.0 mmoles), and BH_3 (12.1 mmoles) in tetrahydrofuran,[7] was added at room temperature to a freshly prepared solution of ethyl (dimethylsulfuranylidene)acetate, prepared[6] from 12.2 g (53.4 mmoles) of (carbethoxymethyl)dimethylsulfonium bromide, and 1.16 g (48.5 mmoles) of sodium hydride in 200 ml of tetrahydrofuran.

After being stirred for 2 hr at room temperature, the solution was refluxed for an additional 3 hr. The solution was filtered and was treated with 10.0 ml of 3 *N* sodium hydroxide solution and 10.0 ml of 30% hydrogen peroxide at 30–40°. After being stirred for 1 hr at room temperature, the solution was saturated with potassium carbonate and extracted with three 50-ml portions of tetrahydrofuran. At this point, a 52% yield of ethyl *n*-nonanoate was indicated by glpc analysis.

The solvent was removed after the solution was dried over anhydrous magnesium sulfate, and the residue was distilled to afford 2.99 g, 45%, of ethyl *n*-nonanoate, bp 108–109° (8 mm), identified by comparison with an authentic sample.

The reaction of N,N-diethyl(dimethylsulfuranylidene)acetamide[6] (VI) with organoboranes proceeds in a manner analogous to that discussed above. For example, tri-*n*-hexylborane reacts with an equimolar quantity of VI to afford, after oxidation, 85% of the theoretical yield of the corresponding amide (*i.e.*, VII). Using a 4:1 molar ratio of ylide to organoborane, the over-all yield of VII is *ca.* 50%.

$$(CH_3)_2\overset{+}{S}\overset{-}{C}HCON(C_2H_5)_2 + R_3B \xrightarrow{THF} \xrightarrow[OH^-]{H_2O_2} RCH_2CON(C_2H_5)_2$$

VI VII

The functionalization of boron-bound groups is not restricted to the use of carbonyl-stabilized ylides such as IV and VI. For example, we have shown that equimolar quantities of triphenylarsonium benzylide[11] and tri-*n*-hexylborane react at room temperature under the usual conditions[4] to afford, after oxidation, 1-phenylheptane and 1,2-diphenyloctane, the product of double homologation, in a 3:1 ratio.[12] Thus, the

$$(C_6H_5)_3\overset{+}{A}s\overset{-}{C}HC_6H_5 + R_3B \xrightarrow[OH^-]{H_2O_2} C_6H_5CH_2R + C_6H_5CH_2CH(C_6H_5)R + ROH$$

$$R = n\text{-}C_6H_{13}$$

product benzylboranes undergo hydrolytic cleavage during the course of the oxidative step.

Acknowledgment. We wish to thank the U. S. Army Research Office (Durham) (DA-ARO-31-124-G889) for generous support of this work.

(11) A. W. Johnson and J. O. Martin, *Chem. Ind.* (London), 1726 (1965).

(12) The corresponding triphenylphosphonium benzylide has been shown[3] to react with organoboranes in a similar fashion, albeit under considerably more forcing conditions.

J. J. Tufariello, L. T. C. Lee, P. Wojtkowski

Department of Chemistry, State University of New York at Buffalo Buffalo, New York 14214

Received August 14, 1967

Additions and Corrections

The Stereochemistry of the Coordination Group in an Iron(III) Derivative of Tetraphenylporphine [*J. Am. Chem. Soc.*, **89**, 1992 (1967)]. By J. L. HOARD, G. H. COHEN, AND M. D. GLICK. Department of Chemistry, Cornell University, Ithaca, New York 14850.

On page 1994 in Table I, the entries under B_x and B_y (columns 5 and 6) for Cl (line 2) should both read 5.4.

One-Step Synthesis of Bridged Aziridines [*J. Am. Chem. Soc.*, **89**, 5045 (1967)]. By WATARU NAGATA, SHOICHI HIRAI, KYOZO KAWATA, AND TSUTOMU AOKI. Shionogi Research Laboratory, Shionogi & Co., Ltd., Fukushima-ku, Osaka, Japan.

On page 5045, column 2, in line 11, $C_{22}H_{35}O_3N$ should read $C_{22}H_{33}O_3N$; in line 16, $C_{20}H_{29}O_3N$ should read $C_{20}H_{27}O_3N$; in line 19, $C_{20}H_{29}O_3N$ should read $C_{20}H_{27}O_3N$; and in line 24, $C_{22}H_{34}O_3NI \cdot 0.5H_2O$ should read $C_{22}H_{32}O_3NI \cdot 0.5H_2O$.

A New Method for Isoquinuclidine Synthesis. A Total Synthesis of Desethylibogamine [*J. Am. Chem. Soc.*, **89**, 5046 (1967)]. By WATARU NAGATA, SHOICHI HIRAI, KYOZO KAWATA, AND TAMOTSU OKUMURA. Shionogi Research Laboratory, Shionogi & Co., Ltd., Fukushima-ku, Osaka, Japan.

On page 5047, column 2, in line 9, $C_{17}H_{22}N_2O_2$ should read $C_{17}H_{20}N_2O_2$.

Stereoselectivity in Hydrogen Atom Transfer to the Vinyl Radicals Derived from the *cis*- and *trans*-*t*-Butyl α-Chloropercinnamates [*J. Am. Chem. Soc.*, **89**, 5251 (1967)]. By LAWRENCE A. SINGER AND NOLAND P. KONG. Department of Chemistry, University of Chicago, Chicago, Illinois 60639.

In footnote 15, we incorrectly cited the olefin distribution obtained from the thermal decomposition of either *cis*- or *trans*-*t*-butyl α-methylpercinnamate in cumene as *trans*:*cis* = 60:40. It should read *trans*:*cis* = 40:60 [L. A. Singer and N. P. Kong, *J. Am. Chem. Soc.*, **88**, 5213 (1966)]. This error invalidates the argument (on page 5254) which claims that the similarity of this ratio to that obtained from the *cis*-α-chloro system is a means of deciding that the product ratios obtained from the *cis*-α-chloro perester were more nearly the expected ones whereas those from the *trans*-α-chloro perester were richer in *cis* olefin than expected. This argument presumes that the stereoselectivity features of the α-chloro-β-phenylvinyl radical and the α-methyl-β-phenylvinyl radical are the same. In retrospect, the stereoselectivity features of the hydro-